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Full Length Article

Optimization of natural gas treatment for the removal of CO₂ and H₂S in a novel alkaline-DEA hybrid scrubber[☆]Samuel Eshorame Sanni^{a,*}, Oluranti Agboola^a, Omololu Fagbiele^a, Esther Ojima Yusuf^a, Moses Eterigho Emeteri^b^aDepartment of Chemical Engineering, Covenant University, P.M.B 9221, Ota, Ogun State, Nigeria^bDepartment of Physics, Covenant University, P.M.B 9221, Ota, Ogun State, Nigeria

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ABSTRACT

Contaminated natural gas when carelessly handled, often poses human and equipment related problems ranging from lung and skin infections to corrosion, equipment fouling/failure and reduction in gas quality owing to the presence of acid gases. In this work, four natural gas (NG) samples were treated to remove CO₂ and H₂S using 10–50% Di-ethanolamine (DEA) solutions mixed with 5% w/w 0.1 M calcium hydroxide. The treatment process gave increased acid gas removal at increased DEA concentrations. Based on the simulation results, cost effective treatment of the gas samples, require 0.1 M Ca(OH)₂ and DEA mixed solutions in the range of 27.4–30%. The optimum mixture concentration for the gas treatment was found to be 30% Ca(OH)₂-DEA hybrid solution with feed gas flow rate of 830 km³/h. In terms of pressure energy consumption, pumping the hybrid mix at 830 km³/h will save pressure energy as compared to pumping the feed gas at 1024.58 km³/h since the lower and upper limit feed gas flow rates gave similar results. The optimum pressure for NG treatment was found to be in the range of 2–2.7 bar (2–2.7*10⁵ kgm⁻² s⁻²).

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1. Introduction

The risks posed by CO₂, H₂S and other impurities in Natural Gas (NG) are increasingly alarming owing to their negative consequences in humans, equipment and the environment. Natural gas may be classified as sweet or sour. It comprises of methane, ethane, propane, isobutane, n-butane, nitrogen, CO₂, O₂, isopentane, n-pentane, hexane and H₂ [1]. Other gases include helium, hydrogen sulphide and mercaptans which give the gas its characteristic odour. CO₂ and H₂S are the major pollutants in NG. For easy/hale transportation, <50 ppm CO₂ is desired [2]. CO₂ and H₂S can be trapped using amine solvents, absorption equipment and membranes [3]. Other contaminants, such as carbonyl sulphide, mercaptans, ethane, pentane etc., are usually removed via distillation and absorption. Gas purification depends on the target-solute solubility, partial vapour pressures of the constituents and the spent heat during solvent recovery [4,5]. Absorption of CO₂ from flue gas using several alcohol-amine solvents has been reported [6].

According to Fang and Zhu [7], the use of amines, carbonates, aqueous ammonia, polymer membranes, ionic liquids and enzymes are recent advances in gas treatment operations. Acid gas removal from contaminated gas depends on the desired application; in internal combustion engines, a CH₄ concentration >90% is friendly [8,9]. High CO₂ in NG reduces engine power [10,11], while >3500 ppm H₂S in gaseous fuels may cause internal corrosion of engines [12]. High amount of CO₂ also reduces the burn-rate of natural gas [3]. Transportation of significant amounts of H₂S can cause pipeline corrosion [13], leakages, fire explosions and loss of aquatic and human lives [14] hence, the gas must be treated prior transportation [15]. In selecting a solvent for gas treatment, one must consider its solubility, viscosity, solvent corrosivity, density, thermal stability, % H₂S/CO₂ in feed gas, process economics and solvent recovery [16,17]. According to reports from National Energy Laboratory, the recommended residual sulphur threshold in NG is 0.3 ppm [18]. One popular amine for acid gas absorption is Di-ethanolamine (DEA) with formula HO(CH₂CH₂OH)₂ [19,20]. Natural gas contains significant quantities of H₂S and CO₂ [21] thus, controlling these contaminants is critical in quantifying the risks associated with equipment fouling. Natural gas liquefies at <161 °C and 1 atm which necessitates CO₂ removal [22,23].

Fig. 1 illustrates a traditional gas treatment process (i.e. the Rectisol process) where cold CH₃OH is used to absorb acid gases from

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